Copy of a printed diagram referenced as "On element of the diamond crystal"

Contributors

Fuller, Watson, 1935-

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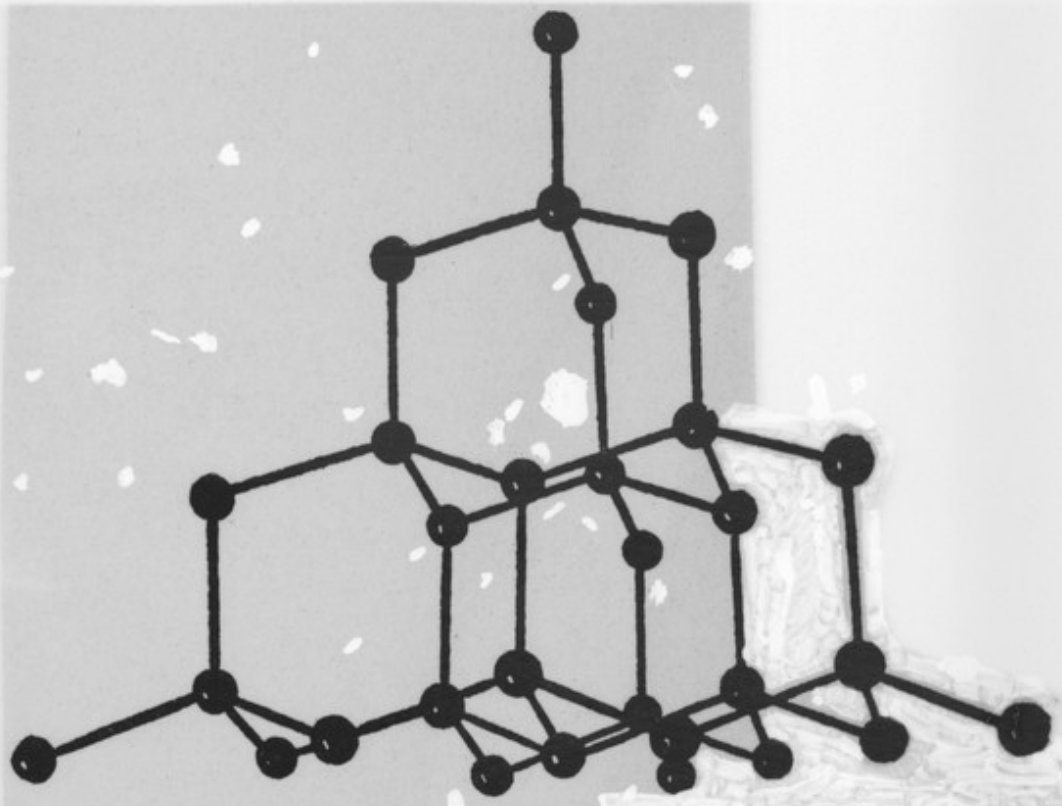
lid, to be contrasted [redacted] om and constant
ing association in the liquid.

Covalent crystals

second class of crystals is termed covalent, because the
that operate to maintain the structure are almost identical
ose already explained for covalent binding. If, for example



An element of the diamond crystal, in which each carbon atom is surrounded tetrahedrally by four others.

e told that a carbon atom, when hybridized in sp^3 style
orm four strong bonds at a tetrahedral angle, we must no
rprised that carbon atoms arrange themselves in an infi
hedral pattern, such as diamond (Fig. 3). It is obvious tha
bond is effectively localized, just as in methane or a large
fin chain. Such solids are really little more than huge mol