**Copy of a printed diagram referenced as "A) The O-hybrids of the carbon atoms of benzene. B) π [pi] -atomic orbitals in benzene and the Kekule pairing schemes"**


**Contributors**

Fuller, Watson, 1935-

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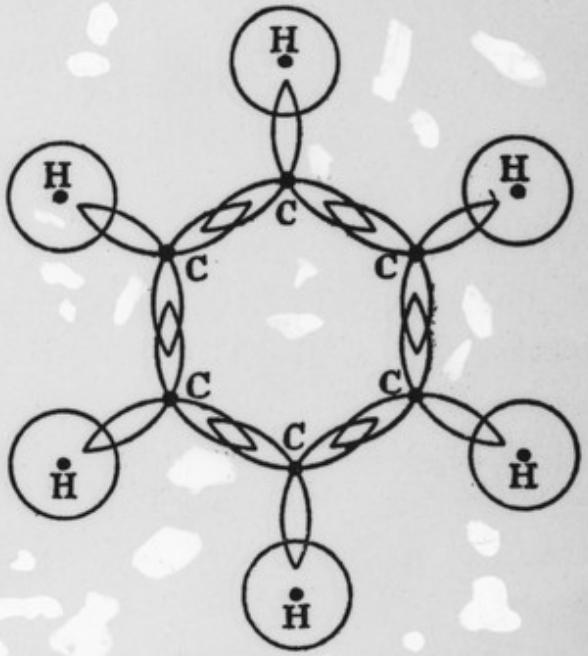
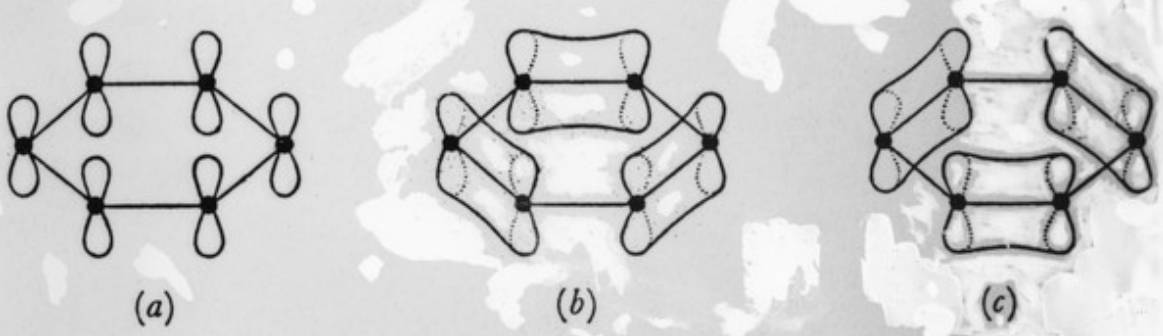
first, for historically it is much the older. There are obviously two quite distinct ways in which we might pair the $\pi$ orbitals in Fig. 2 ($a$). They are shown in Figs. 2 ($b$) and 2 ($c$). Neither of these systems which are 'structures' in the sense of § 5.5 has any




The $\sigma$-hybrids of the carbon atoms of benzene.




The $\pi$-atomic orbitals in benzene ($a$), and the Kekulé pairing schemes ($b$, $c$).

greater merit than the other. Both must therefore appear in the complete wave function. Physically we may say that the benzene molecule behaves partly as if the bonds were in the one position and partly as if in the other.

This is a mathematical statement about the Schrödinger wave