

Printed graph referenced as "absorption spectrum of methyl alcohol in carbon tetrachloride solutions"

Contributors

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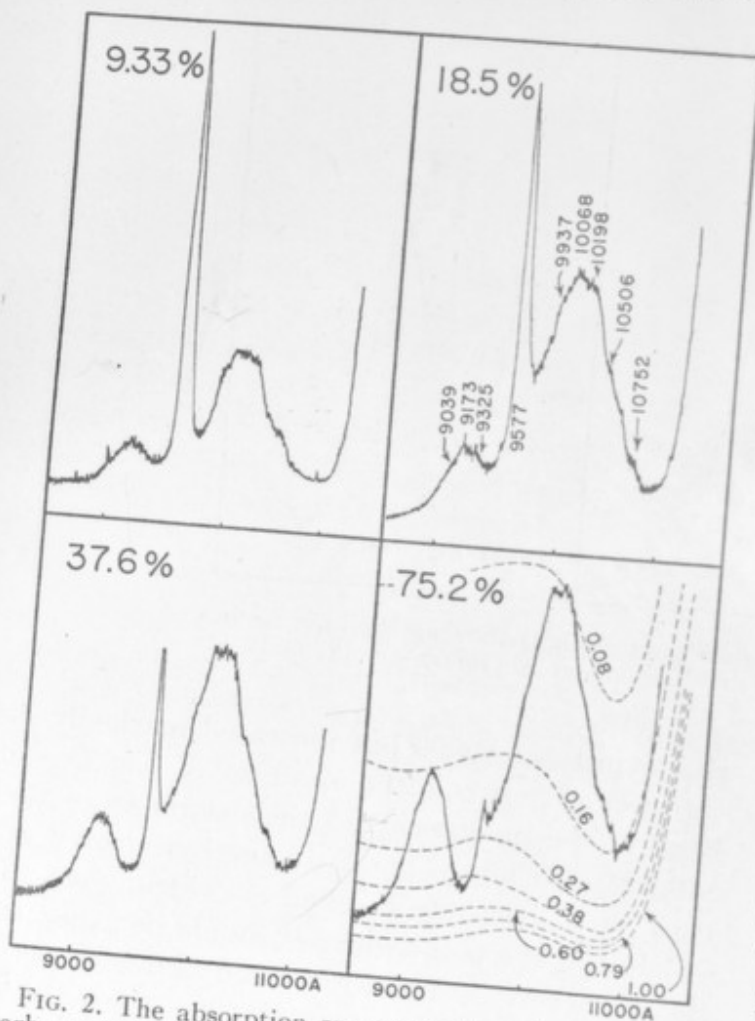


FIG. 2. The absorption spectrum of methyl alcohol in carbon tetrachloride solutions. The concentrations are in mole percent and the path lengths were inversely proportional to the concentration so that the amount of alcohol in the path was constant. The broken curves in the last diagram represent the background blackening with only carbon tetrachloride in the cell and with screens of the transmissivities indicated interposed in the path.

which are designated with wave-lengths are definitely real.

Though the origin of the high frequency association band may be in some doubt we have no hesitation in ascribing the broader and more intense one to a shifted O—H vibration.

It is of association bands simultaneously concentrated from the same source in their temperature in ethyl alcohol.

Broad absorption bands at $\lambda 10,000$ are higher alcohol. Examples are given in the literature on the O—H bands. In the present case the C—H bands are weak in methyl alcohol at $\lambda 9098$, and quite intense. The association bands are distinct from the compounds in the compounds. Some observations will be described in the present, superimposed.

In the aromatic region there are very unfavorable bands. There are bands which do not involve the C=C bond. The result from the C=C vibration.

In the lower frequency bands are practically the same in the liquids, but as the concentration increases they shift to higher frequency. Note for example in Fig. 3. This is